Chapter 16, Objectives

You should be able to:

1. write molecular, complete ionic, and net ionic reaction equations for precipitate reactions,
2. write the Ksp expressions for salts,
3. indicate the significance of a large Ksp and a small Ksp,
4. use the Ksp expression to determine the concentration of ions at equilibrium,
5. use the Ksp to calculate water solubility in moles or grams per volume of solution,
6. work problems involving the common ion effect,
7. use Ksp to calculate the concentration of ion required to start a precipitation,
8. use Ksp to predict whether or not a precipitate will form,
9. determine which precipitate will form first during a fractional precipitation and determine the concentration of ion required to start a precipitation,
10. predict whether or not a precipitate will dissolve in an acidic solution and write the equation for the reaction that would occur.